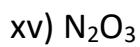
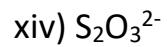
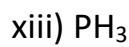
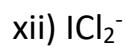
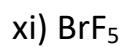
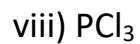
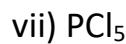
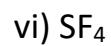
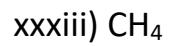
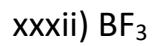
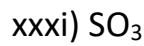
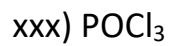
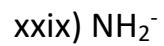
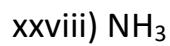
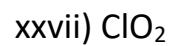
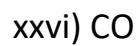
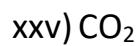
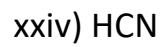
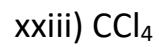
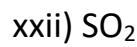
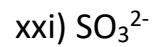
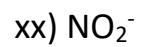
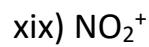
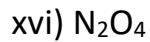


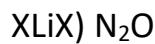
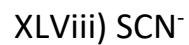
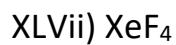
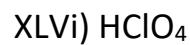
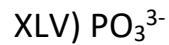
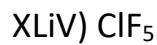
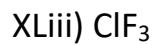
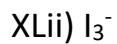
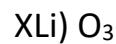
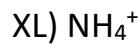
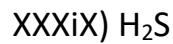
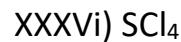
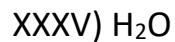
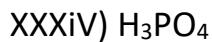


UNIT 2: STRUCTURE AND BONDING(1)

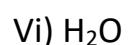
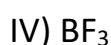
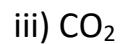
1) Draw the Lewis dot dash diagrams for the following molecules.

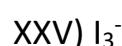
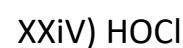
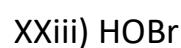
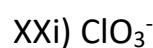
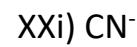
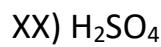
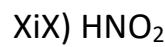
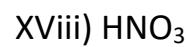
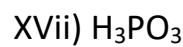
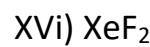
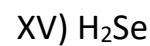
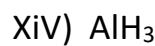
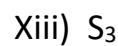
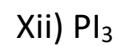
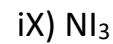


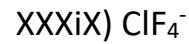
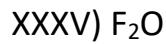
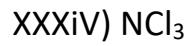
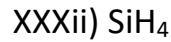
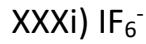
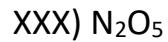
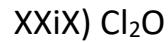
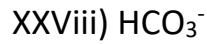




2) Draw the Lewis dot-dash diagrams for the following molecules.







3) Predict the shapes of the following molecules using the VSEPR theory.



iv. BCl_3

v. AlCl_3

vi. BH_3

vii. AlF_3

viii. CCl_4

ix. CHCl_3

x. CH_2Cl_2

xi. CH_3Cl

xii. SiH_4

xiii. SiCl_4

xiv. PCl_5

xv. SF_6

xvi. IF_7

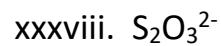
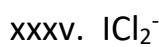
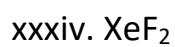
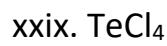
xvii. H_2O

xviii. NO_2

xix. NO_3^-

xx. NO_2^-

xxi. SO_4^{2-}



XL. POClBrI

XLi. ClO₂⁻

XLii. ClO₃⁻

XLiii. ClO₄⁻

XLiV. NH₄⁺

XLV. BrF₅

XLVi. CO₃²⁻

XLVii. CO₂

XLViii. SO₂

XLix. O₃

L. XeF₆

Li. XeO₃

Li. BF₄⁻

Liii. H₃O⁺

LiV. HCN

LVi. H₂S

LVii. ICl₄⁻

LViii. IF₅